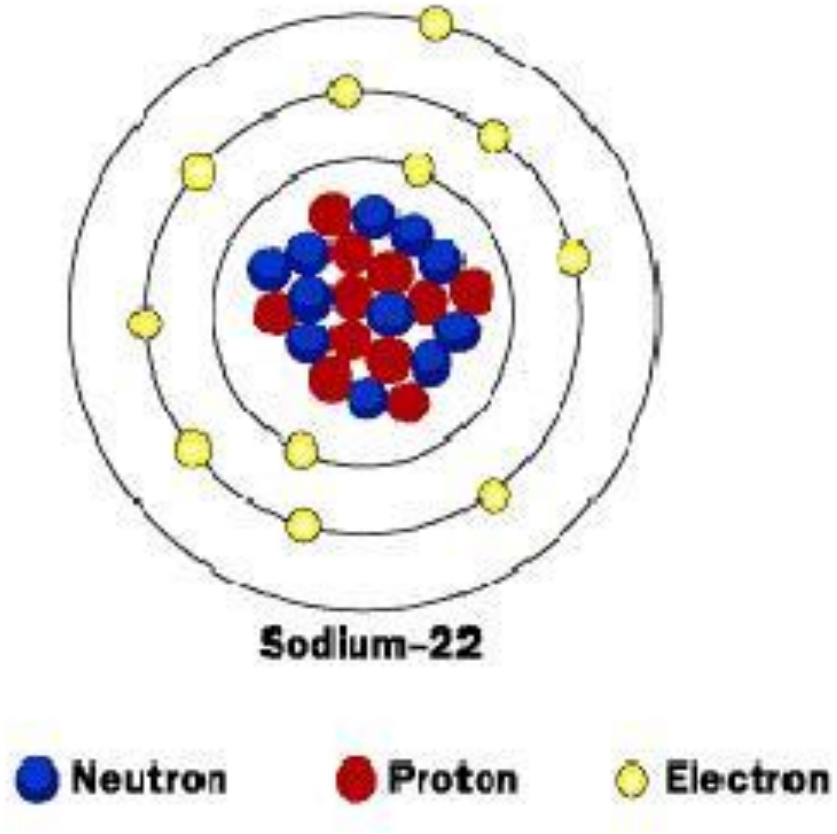
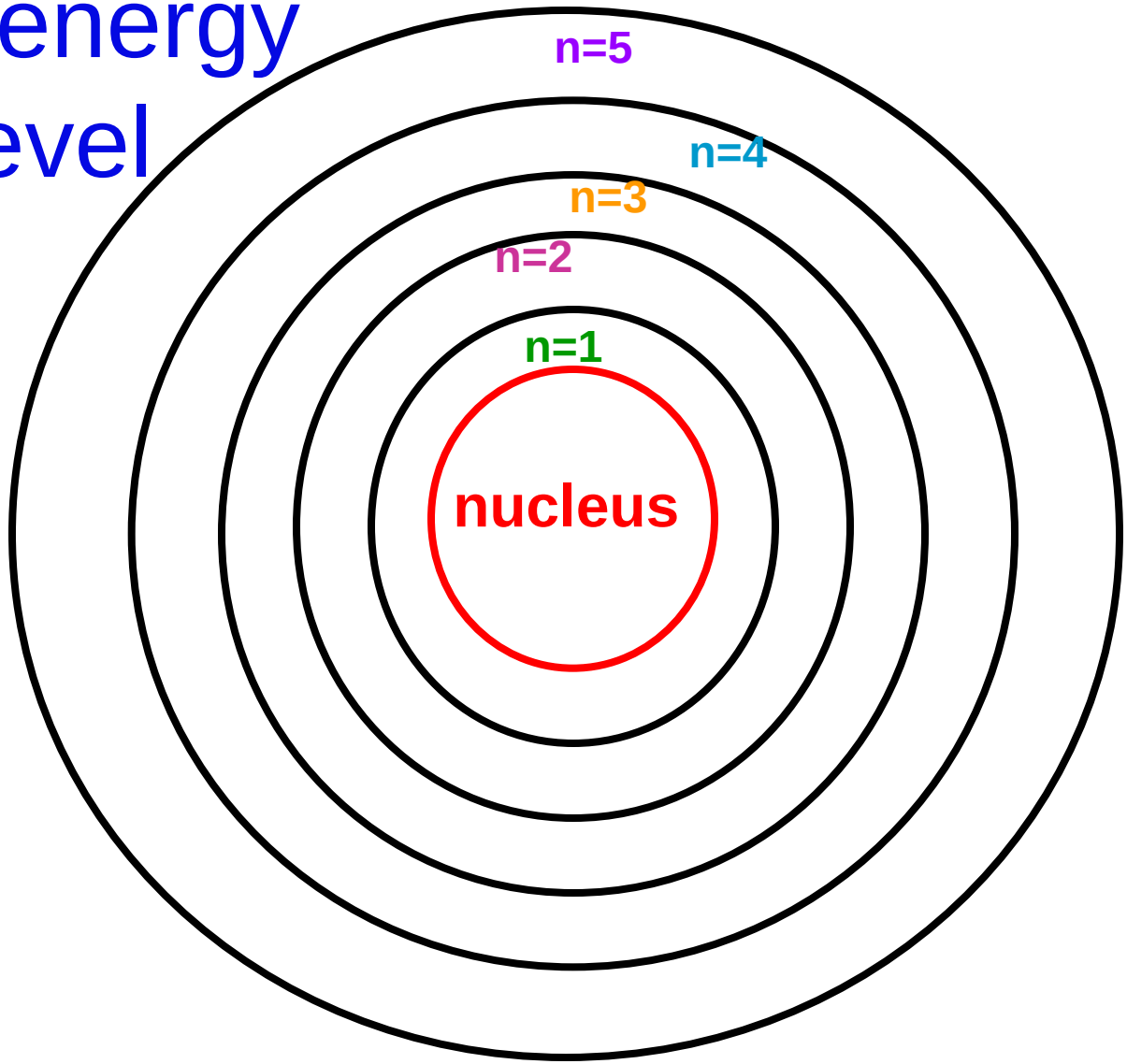
Understanding the Bohr Model



of e^- = $2(n^2)$

n = energy level



n=1

e=2

n=2

e=8

n=3

e=18

n=4

e=32

n=5

e=50

of e^- = $2(n^2)$

n = energy level

Oxygen O

Mass number = 16

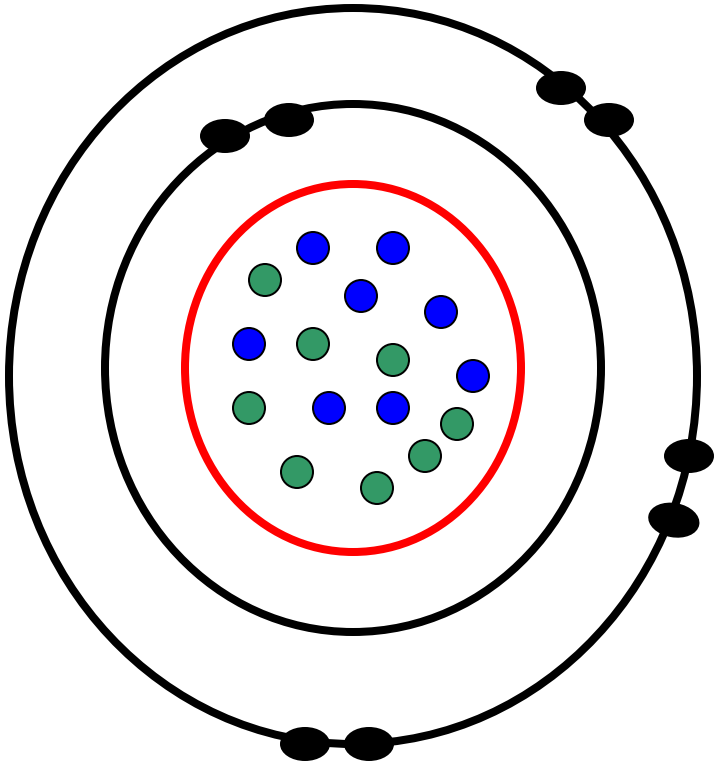
of p^+ = 8 [from the atomic number]

of e^- = 8 [from the # of protons]

of $n^\pm$ = 8 [mass # minus protons]

Now we can make our Bohr model.

- = e^- = 8
- = $n^{\pm}$ = 8
- = p^+ = 8



of e^- = $2(n^2)$

n = energy level

Mercury Hg

Mass number = 201

of p^+ = 80

of e^- = 80

of $n^\pm$ = 121

Now we can make our Bohr model.

Short Cut Method for Hg

$80e^-$

$- 2e^-$

$78e^-$

$- 8e^-$

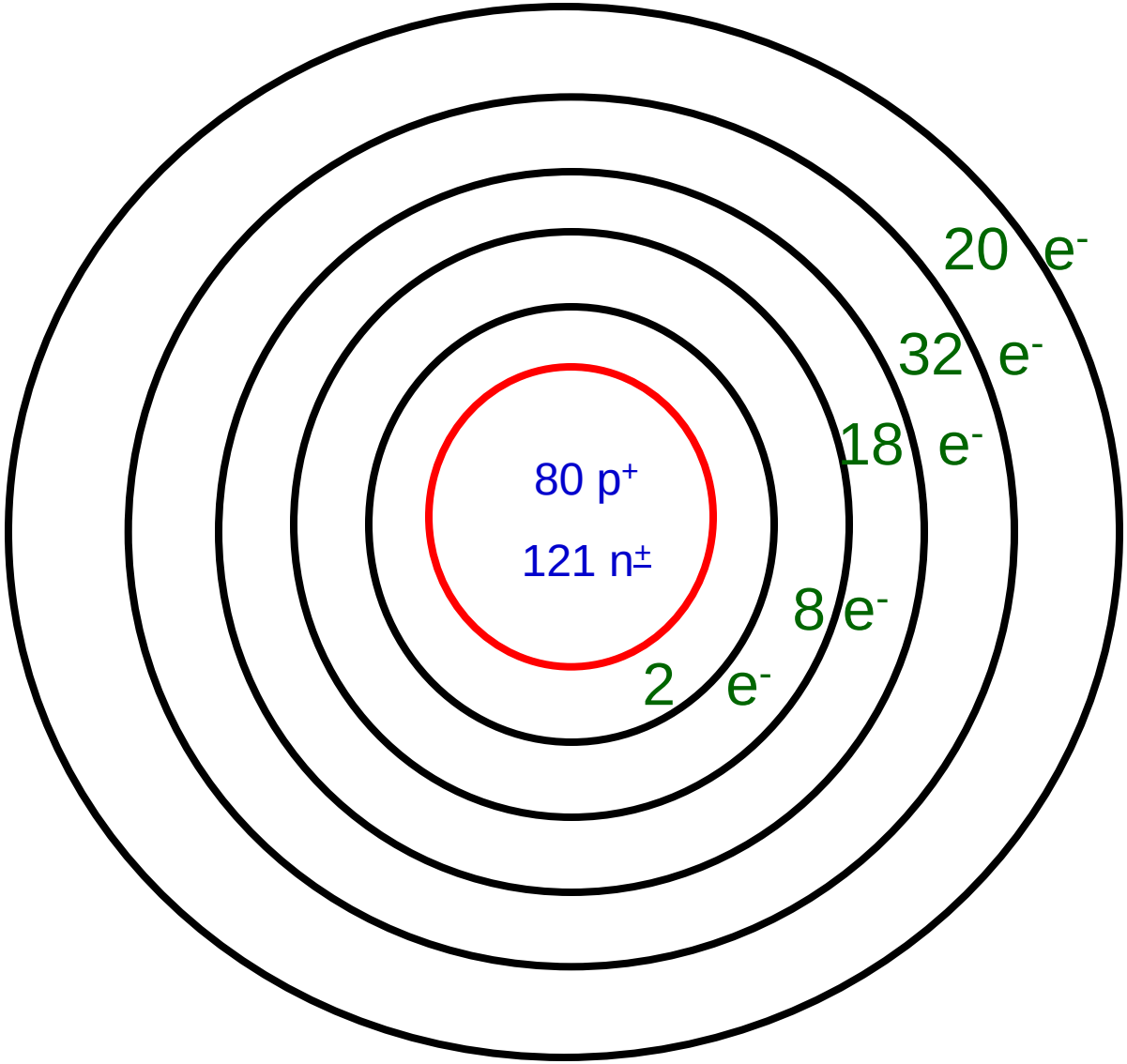
$70e^-$

$-18e^-$

$52e^-$

$- 32e^-$

$20e^-$



of e^- = $2(n^2)$

n = energy level

Bromine Br

Mass number = 80

of p^+ = 35

of e^- = 35

of $n^\pm$ = 45

Now we can make our Bohr model.

Short Cut Method for Br

$35e^-$

$-\ 2e^-$

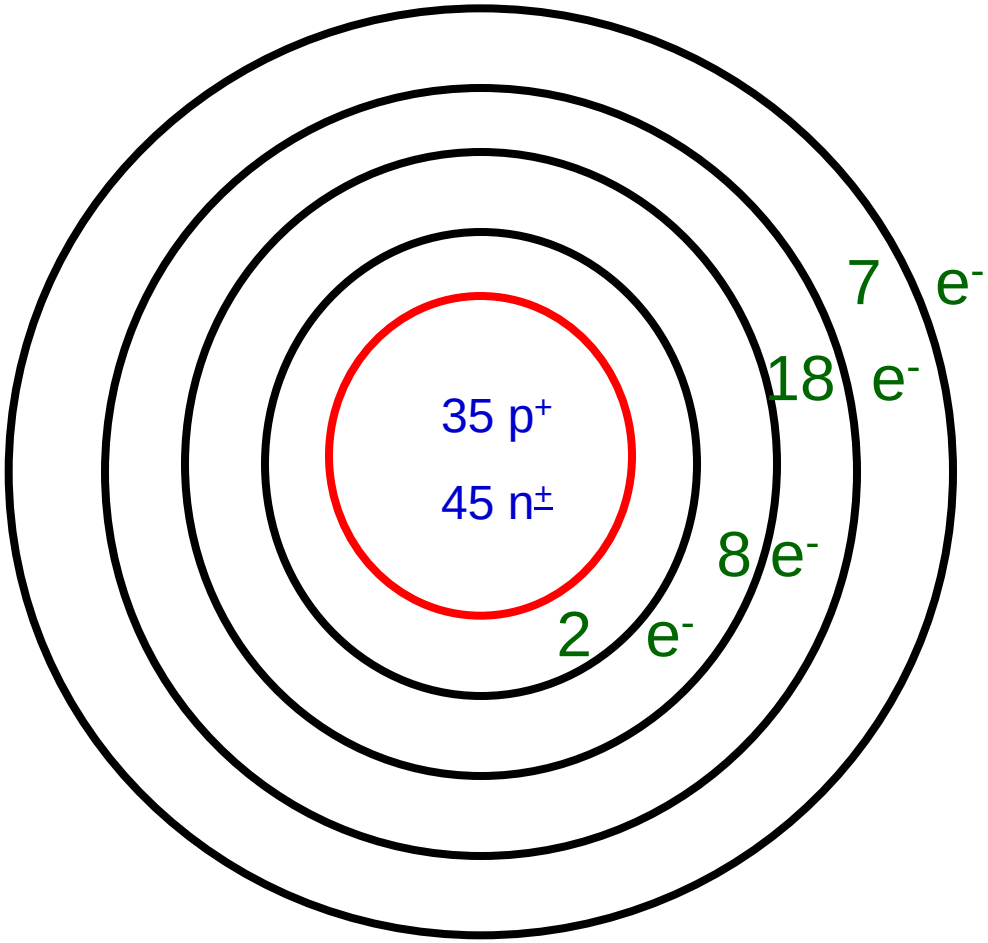
$33e^-$

$-\ 8e^-$

$25e^-$

$-18e^-$

$7e^-$



of e^- = $2(n^2)$

n = energy level

Xenon Xe

Mass number = 131

of p^+ = 54

of e^- = 54

of $n^\pm$ = 77

Now we can make our Bohr model.

Short Cut Method for Xe

$54e^-$

$- 2e^-$

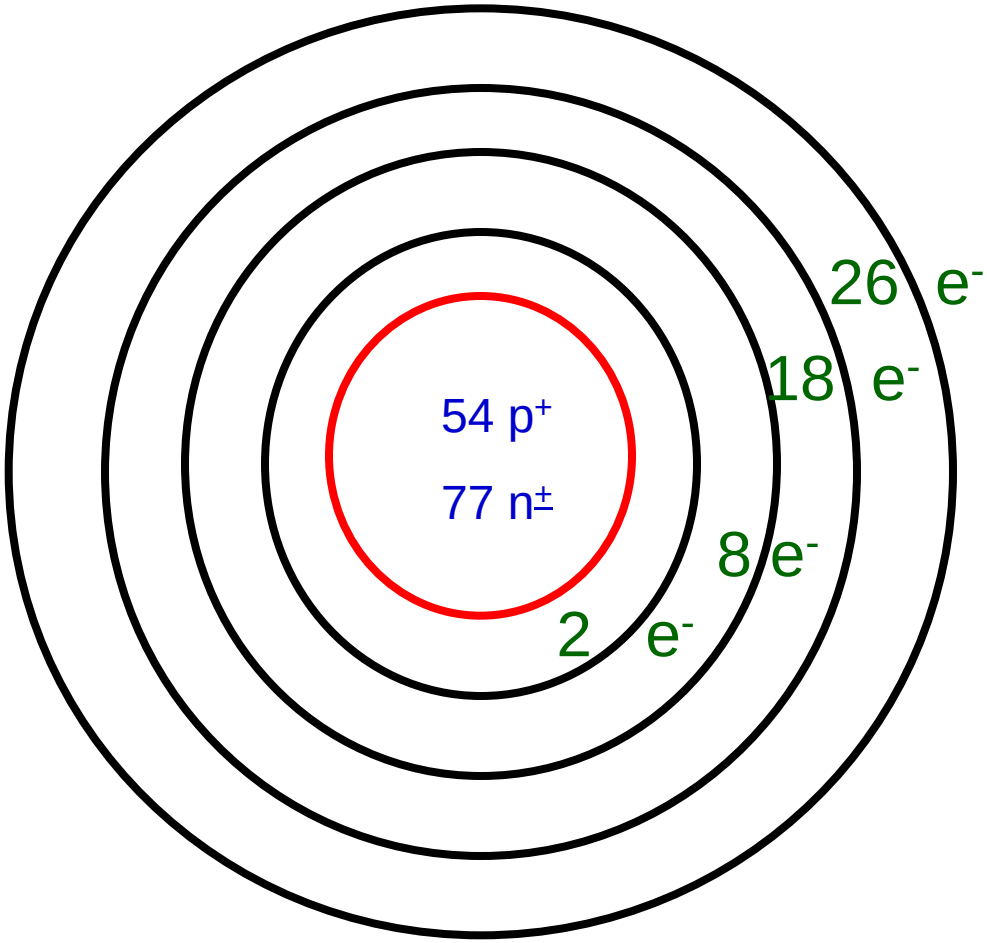
$52e^-$

$- 8e^-$

$44e^-$

$-18e^-$

$26e^-$



Website Check

- Go to my website to check your Bohr model and find element information.

http://www.chemicalelements.com/

USE YOUR RUBRIC!!!! Due Monday!!

Lost your rubric copy it from someone else!